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Porphyrins are the conjugate acids of ligands that bind metals to form complexes. The metal ion usually has a charge of 2+ or 3+. A schematic equation for these syntheses is shown: $H_2 \text{ porphyrin} + [ML_n]^{2+} \rightarrow M(\text{porphyrinate})L_n + 4L + 2H^+$, where M = metal ion and L = a ligand. A porphyrin without a metal-ion in its cavity is a free base. Some iron-containing porphyrins are called hemes.

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Exp. Pharmacol. Handbook of Experimental Pharmacology Hazard. Waste Hazard. Mater ... A Journal of Photochemistry and Photobiology A J. Photochem. Photobiol., B Journal of ... Chem Jpn. Journal of Synthetic Organic Chemistry, Japan J. Test. Eval. Journal of Testing and Evaluation J . Therm ...

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ADDRESSED IN THIS GUIDANCE MANUAL 1-8 1.6 A SUMMARY OF ALTERNATIVE DISINFECTANT PROPERTIES 1-9 1.7 SELECTING A ...

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